

Monovalent Metals

Lithium - martensitic transformation to a mixture of crystalline phases 77K

Sodium - at 23K, ok

H metallic @ high pressure \square

Alkali - BCC

Noble - FCC

Alkali BCC: $\frac{k_F^3}{3\pi^2} = n = \frac{2}{a^3}$

$$k_F = 0.62 \left(\frac{2\pi}{a} \right)$$

In

Reciprocal: FCC

$$k_{min} = \frac{4\pi}{a} \sqrt{\left(\frac{1}{2}\right)^2 + \left(\frac{1}{2}\right)^2} = \frac{4\pi}{a} \sqrt{\frac{1}{2}}$$

$$\frac{k_{min}}{2} = 0.707 \left(\frac{2\pi}{a} \right)$$

dHvA very nearly sphere

Hall $-1/nec$ (few%)

Noble Metals: Filled d

6 bands

characterize as s, d

s dep. like free el.

asymmetry magnetoresistance

Optical Properties (Appendix K)

2.

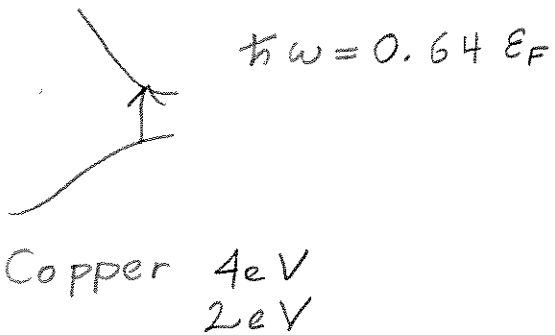
Color freq. dep. of reflectivity (some more others)

Plasmons

Copper ω_p : red
gold yellow
silver blue

Tanner?

$$k' = k + q + K, \quad k_F \approx 10^8 \text{ cm}^{-1}$$
$$q \approx 10^5 \text{ cm}^{-1} \approx 0$$
$$k' = k$$



Divalent Metals

3.

d not as important

Ca, Sr, Ba ^{free el.} sphere same vol. first zone

good dHvA Be, Mg, Zn, Cd

Be compensated

Trivalent

Only consider Al ↙ very small

Free EL. 2nd 3rd 4th BZ

2nd zone empty

3rd zone narrow tubes occupied

$$n_e^{\text{II}} + n_e^{\text{III}} = \frac{n}{3}$$

$$n_e^{\text{II}} + n_h^{\text{II}} = 2\left(\frac{n}{3}\right)$$

$$n_e^{\text{III}} - n_h^{\text{II}} = -\frac{n}{3} \quad R_H = \frac{-1}{nc} \quad \text{pos. \& dens} = \frac{1}{3} \text{ of } n.$$

absorb. \& refl. dHvA

Tetravalent

Lead \& Tin

Lead compare to Al

$$n_h^{\text{II}} = n_e^{\text{III}} \quad \text{derive?}$$

Semimetals

$$n_v \ll 10^{22}/\text{cm}^3$$

$$\text{Graphite } n_e = n_h = 3 \times 10^{18}/\text{cm}^3$$

Pentavalent metals: As, Sb, Bi

- > Transition metals (d)
- > Rare earths (f) localized
- > Alloys Stoichiometric... basis
Disorder ...elaborate

NIST, text, FS elements

Nd Fe B
Sm Co

Poole Cu, Ag, Au 10.8, 9.0, 9.0 eV

www.ee.iitm.ac.in/~hsr/ec301/copper.pdf

INDIA

Atomic Properties of the Elements

PERIODIC TABLE

18
VIII A

NIST
National Institute of
Standards and Technology
U.S. Department of Commerce

Physical Measurement Laboratory

www.nist.gov/ptl

Reference Data

www.nist.gov/std

Frequently used fundamental physical constants

For the most accurate values of these and other constants, visit physics.nist.gov/constants

1 second = 9 192 631 770 periods of radiation corresponding to the transition between the two hyperfine levels of the ground state of ¹³³Cs

Planck constant h

elementary charge e

electron mass m_e

proton mass m_p

fine-structure constant α

Rydberg constant R_∞

Boltzmann constant k

$299\,792\,458\text{ m s}^{-1}$ (exact)
 $6.626\,07 \times 10^{-34}\text{ J s}$ (exact)
 $1.602\,177 \times 10^{-19}\text{ C}$
 $9.108\,38 \times 10^{-31}\text{ kg}$
 $0.510\,989\text{ MeV}$
 $1.672\,622 \times 10^{-27}\text{ kg}$
 $1/137.035\,999$
 $10\,973\,731.569\text{ m}^{-1}$
 $3.289\,841\,960 \times 10^{16}\text{ Hz}$
 $1.380\,658 \times 10^{-23}\text{ J K}^{-1}$

- Solids
- Liquids
- Gases
- Artificially Prepared

Group	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18																	
Period	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18																	
1	H Hydrogen 1.008 1s ¹	He Helium 4.002602 1s ²											Li Lithium 6.94 2s ¹	Be Beryllium 9.012182 2s ²	B Boron 10.81 2s ² 2p ¹	C Carbon 12.011 2s ² 2p ²	N Nitrogen 14.007 2s ² 2p ³	O Oxygen 15.999 2s ² 2p ⁴	F Fluorine 18.9984032 2s ² 2p ⁵	Ne Neon 20.1797 2s ² 2p ⁶															
2	Na Sodium 22.98976928 3s ¹	Mg Magnesium 24.3050 3s ²	Al Aluminum 26.9815386 3s ² 3p ¹	Si Silicon 28.085 3s ² 3p ²	P Phosphorus 30.973762 3s ² 3p ³	S Sulfur 32.06 3s ² 3p ⁴	Cl Chlorine 35.45 3s ² 3p ⁵	Ar Argon 39.948 3s ² 3p ⁶	K Potassium 39.0983 4s ¹	Ca Calcium 40.078 4s ²	Sc Scandium 44.955912 3d ¹ 4s ²	Ti Titanium 47.867 3d ² 4s ²	V Vanadium 50.9415 3d ³ 4s ²	Cr Chromium 51.9961 3d ⁵ 4s ¹	Mn Manganese 54.938045 3d ⁵ 4s ²	Fe Iron 55.845 3d ⁶ 4s ²	Co Cobalt 58.933195 3d ⁷ 4s ²	Ni Nickel 58.6934 3d ⁸ 4s ²	Cu Copper 63.546 3d ¹⁰ 4s ¹	Zn Zinc 65.38 3d ¹⁰ 4s ²	Ga Gallium 69.723 4s ² 4p ¹	Ge Germanium 72.63 4s ² 4p ²	As Arsenic 74.92160 4s ² 4p ³	Se Selenium 78.96 4s ² 4p ⁴	Br Bromine 79.904 4s ² 4p ⁵	Kr Krypton 83.798 4s ² 4p ⁶									
3	Rb Rubidium 85.4678 5s ¹	Sr Strontium 87.62 5s ²	Y Yttrium 88.90585 4d ¹ 5s ²	Zr Zirconium 91.224 4d ² 5s ²	Nb Niobium 92.90638 4d ⁴ 5s ¹	Mo Molybdenum 95.96 4d ⁵ 5s ¹	Tc Technetium (98) 4d ⁵ 5s ²	Ru Ruthenium 101.07 4d ⁷ 5s ¹	Rh Rhodium 102.90550 4d ⁸ 5s ¹	Pd Palladium 106.42 4d ¹⁰	Ag Silver 107.8682 4d ¹⁰ 5s ¹	Cd Cadmium 112.411 4d ¹⁰ 5s ²	In Indium 114.818 5s ² 5p ¹	Sn Tin 118.710 5s ² 5p ²	Sb Antimony 121.760 5s ² 5p ³	Te Tellurium 127.60 5s ² 5p ⁴	I Iodine 126.90447 5s ² 5p ⁵	Xe Xenon 131.29 5s ² 5p ⁶	Cs Cesium 132.9054518 6s ¹	Ba Barium 137.327 6s ²	La Lanthanum 138.90547 5d ¹ 6s ²	Ce Cerium 140.116 5d ¹ 6s ²	Pr Praseodymium 140.90765 5d ¹ 6s ²	Nd Neodymium 144.242 5d ¹ 6s ²	Pm Promethium (145) 5d ¹ 6s ²	Sm Samarium 150.36 5d ¹ 6s ²	Eu Europium 151.964 5d ¹ 6s ²	Gd Gadolinium 157.25 5d ¹ 6s ²	Tb Terbium 158.92535 5d ¹ 6s ²	Dy Dysprosium 162.500 5d ¹ 6s ²	Ho Holmium 164.93032 5d ¹ 6s ²	Er Erbium 167.259 5d ¹ 6s ²	Tm Thulium 168.93421 5d ¹ 6s ²	Yb Ytterbium 173.054 5d ¹ 6s ²	Lu Lutetium 174.967 5d ¹ 6s ²
4	Fr Francium (223) 7s ¹	Ra Radium (226) 7s ²	Ac Actinium (227) 6d ¹ 7s ²	Th Thorium 232.03806 6d ² 7s ²	Pa Protactinium 231.03688 5f ² 6d ¹ 7s ²	U Uranium 238.02891 5f ³ 6d ¹ 7s ²	Np Neptunium (237) 5f ⁴ 6d ¹ 7s ²	Pu Plutonium (244) 5f ⁶ 6d ¹ 7s ²	Am Americium (243) 5f ⁷ 7s ²	Cm Curium (247) 5f ⁷ 6d ¹ 7s ²	Bk Berkelium (247) 5f ⁷ 6d ¹ 7s ²	Cf Californium (251) 5f ¹⁰ 7s ²	Es Einsteinium (252) 5f ¹¹ 7s ²	Fm Fermium (257) 5f ¹² 7s ²	Md Mendelevium (258) 5f ¹³ 7s ²	No Nobelium (259) 5f ¹⁴ 7s ²	Lr Lawrencium (262) 5f ¹⁴ 6d ¹ 7s ²																		
5	13 IIIA	14 IVA	15 VA	16 VIA	17 VIIA	18 VIIIA																													

¹Based upon ¹²C. (¹) indicates the mass number of the longest-lived isotope. ²IUPAC conventional atomic weights; standard atomic weights for these elements are expressed in intervals; see iupac.org for an explanation and values. For a description of the data, visit physics.nist.gov/data (NIST SP 966 (March 2013))